



Bridge between research in modern physics
and entrepreneurship in nanotechnology

Quantum Physics

*The physics of the very small
with great applications*

Part 2

QUANTUM PROPERTIES & TECHNOLOGY



Learning Station IX:

Spin and its applications



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Table of Contents

Part 2: Quantum Properties & Technology

LEARNING STATION IX : SPIN AND ITS APPLICATIONS	3
1 Undefined properties of matter	3
2 How does spin behave in a magnetic field	4
3 Quantum weirdness: consecutive Stern-Gerlach experiments	5
4 Application: Magnetic Resonance Imaging (MRI)	8
5 Application: Spintronics	9
6 Concepts in Learning Station IX	2

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Learning station IX :

Spin and its applications

1 Undefined properties of matter

When you start learning physics you hear about properties of matter that seem familiar to you from daily life: mass, temperature, energy... But are these quantities really as familiar and as clear as you think? Would you be able to explain what temperature or energy really is? It turns out it takes some mental (and experimental) work to learn how reality works. And in science we try to be as accurate and precise about the statements which are being made. Which is why we try to define the words we use as often as possible. Nevertheless, not everything can be defined. We have to start somewhere, with words and ideas that just are what they are. And hope that we all agree on what they mean. It is possible to precisely define concepts like temperature, energy and force.

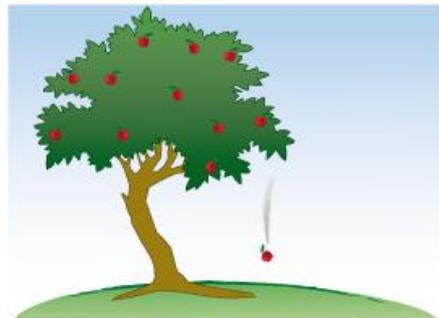
Mass on the other hand is a property that objects (think for instance of a ball or a car) simply seem to have. We do not really know what mass is (however, the recent discovery of the Higgs boson helps us a little further on our way to understand mass). Simply stated, mass is that property of matter which is measured in kilograms.

But we do know how this property of matter manifests itself. Can you think of formulae in physics in which mass plays an important role?

.....

Figure 1: A manifestation of mass?

(source: www.dlswb.rmit.edu.au)



We often tend to relate the mass of an object to the amount of "stuff" present. But is this actually true?

If you take 2 identical buckets and fill one of them with water and the other with oil you will notice when weighing them that oil is lighter. Its mass is smaller. Does this mean there is less oil than water? yes / no

Because of this scientists think of mass in a different way.

They see that mass appears in several well-known laws of physics:

- Newton's second law
- Gravitational force

And through these laws they give meaning to the word mass. It turns out mass plays 3 main roles in physics:

1. The more mass there is, the harder it is to **accelerate** the object
2. The more mass there is, the more it is attracted by the force of **gravity**
3. The more mass there is, the more it creates a **gravitational field** around it

This is not a text about mass. But the main point is that there are *some properties* of materials that just seem to exist and, properties that causes *interaction with surroundings*.

Another such property is **electrical charge** which tells us something about *how the object will interact with electrical and magnetic fields*.

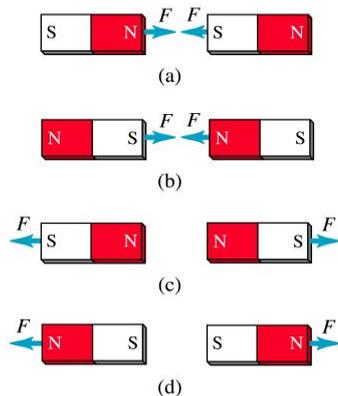
The property that we want to study in this text is called **spin**. It is one more property of matter which cannot be defined which is why we will not be able to tell you what it is! You may complain about that. But just remember that you were happy enough to talk about the mass of an object without anyone ever telling you what it actually is.

But what can be learned about spin is how it affects the behavior of the objects. This is what we are going to discuss in the next sections.

2 How does spin behave in a magnetic field

It turns out that *objects with spin behave like small magnets*. Which means among other things that they will be affected by the presence of other magnets. What happens when you take 2 magnets and bring them close to each other?

.....



Does the result depend on which side of each magnet faces the other?

.....

Every magnet has 2 sides, known as its north and its south pole. The interaction between 2 magnets is shown here on the figure. The strength of the attraction or the repulsion between them depends on the strength of the magnets themselves.

(Source figure: www.physics.sjsu.edu/becker/physics51/mag_field.htm)

In 1922 *Otto Stern and Walther Gerlach* did an experiment, known today as the Stern-Gerlach experiment, to find out *how a magnetic field affects the motion of objects with spin*.

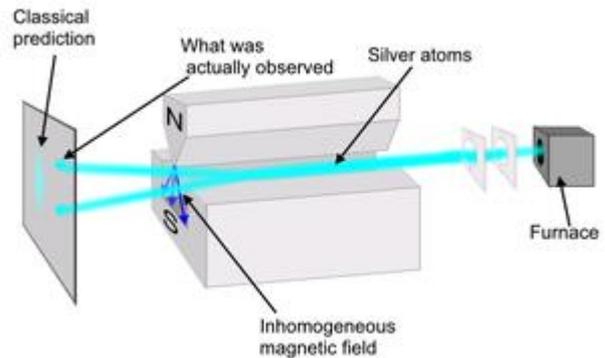
It was known at the time that a *moving electric charge* in a magnetic field would be *deflected*. But even **if the particle had no electric charge it would still be deflected in magnetic field if it had spin**.

They took a furnace from which silver atoms are ejected. These atoms have spin so that they behave like magnets. The beam of silver atoms was passed through a magnetic field created by another magnet. Each silver atom was either attracted towards the north pole of the second magnet (deflecting it upwards in the figure), or repulsed by it (deflecting it downwards). Whether it is attracted or repulsed depends on the orientation of the spin in the silver atom. This means we can separate objects with different spin by letting them go through a magnetic field. Which is exactly what Stern and Gerlach did.

What do you think will determine the amount by which the silver atoms are deflected?

.....

The amount by which it is deflected depends on the value of the spin (the *strength* of the tiny magnet). At first, it was thought that the spin could take *any* value, and so *every possible angle of deflection should be possible*. This would be seen as a vertical line (the blue line on the screen on the left of the figure).



Surprisingly, this is not what actually happened. On the screen there appeared **only 2 dots**. One up, one down. What do we learn about spin from this?

.....

*It seems there are just 2 possible angles of deflection.
There was only 1 possible strength of the tiny magnet.*

*The quantization of the spin was thus discovered. We say the spin is either up or down.
Half the atoms have spin up and half have spin down.*

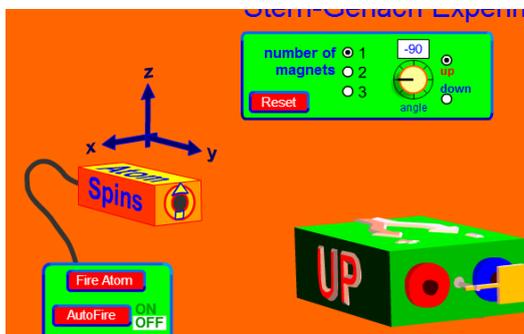
Spin is a quantum-mechanical property, which does not have a counterpart in classical physics.

But this is not the end of the weirdness.

3 Quantum weirdness: consecutive Stern-Gerlach experiments

In reality, objects have more than 1 spin. They have *3 different spins: one in each dimension*. **The atom behaves as if it had 3 perpendicular magnets in it.**

In the figure about the Stern-Gerlach experiment above the magnetic field is pointing vertically and thus separates spin up and spin down in the vertical direction. The experiment could be redone but with a magnetic field pointing from left to right instead.



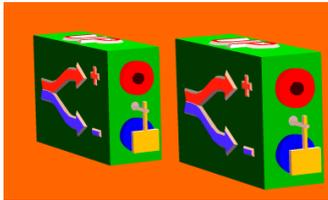
In the phet-applet (see figure hereby) you can rotate the magnetic field through which the beam passes to see what happens.

(Source Phet Colorado

http://phet.colorado.edu/sims/stern-gerlach/stern-gerlach_en.html)

On the left of the figure we have the furnace from which atoms are randomly ejected. On the right of the figure we have the magnet in the green box. The spin up atoms come out of the left hole and the spin down atoms out of the right hole. *We are thus separating spin ups and spin downs in the horizontal direction* (they are still called up and down although

it is known a separation of right and left). If you run the experiment for a certain time you will see that there is about 50% probability that the atom is a spin up one or a spin down one.

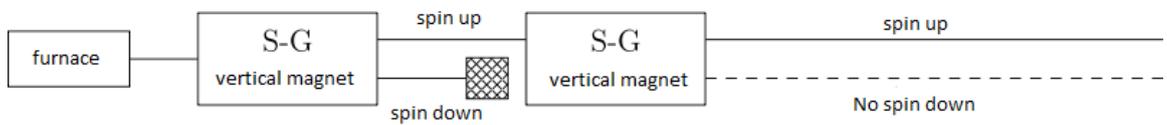


We now take 2 magnets one after the other. Out of the first magnets 2 beams will come out. One with spin up and one with spin down atoms.

If we now put a wall in front of the spin down beam, so that only the spin up atoms would enter the second magnet, what do you expect to see coming out of the second one?

.....

What is actually seen can be depicted by:



Because all the atoms going into the second magnetic field all have spin up it is not surprising that all of them are deflected upwards in the second magnet: there are no spin down atoms present in the beam coming out of the second magnetic field.



Let us now rotate the second magnet by 90 degrees (so that it becomes horizontal). What do you think will happen?

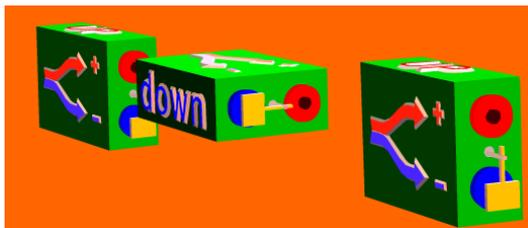
.....

The result turns out to be:



Although the atoms entering this second magnet all have spin up in the vertical direction, their spin in the horizontal direction is still random, which means 50% is spin up and 50% spin down. There are indeed two beams coming out of this second magnetic field.

All of this may not seem too surprising to you, but look at what happens when we add one last magnet, so that it too points in the vertical direction.

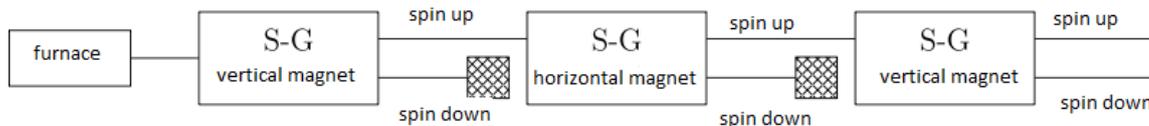


Let us block this time also the spin down coming out of the second magnet. What result would you expect to see?

.....

*Because the **spin down** in the vertical direction was **blocked** after the first magnet, you may expect to **only see spin up atoms** coming out of this third magnet.*

But that is not what actually happens. What is observed are 2 beams coming out of this third magnetic field, *still half of the atoms going up and the other half going down.*



It seems that measuring the spin in the horizontal direction affects the presence of spin up and spin downs in the vertical direction.

This weird effect is a quantum mechanical phenomenon which cannot be understood with classical physics. What happens here reminds us of learning station 4: before making a measurement, the particle is in a "superposition state", in this case of possible states of spin. Only when we make a measurement, the atom (or the particle more in general) appears in a precise spin state, i.e. up or down in a certain direction: we have made a measurement and the "superposition state" is gone away, i.e. the particle is in a well determined state as a consequence of the measurement. That is why there is an equal, 50% probability to get up or down spin.

Now, if we measure the spin along the vertical direction, we determine the state of the spin in that direction. If we make a second measurement in the same direction, then we will obtain the same result, as the state in that direction is now determined. However, when the second measurement determines the state of the spin in another direction, in our example the horizontal one, then the previous information related to the vertical direction is "erased". Making this second measurement in another direction acts in such a way that the superposition of states (i.e. possible spin states) in the initial direction (vertical in the example) is restored!

This quantum mechanical property of spin is not only very interesting, but it turns out to be of vital importance for several applications among which are **medical imaging**, and some aspects of **electronics**. Among many other applications of spin we should not forget to mention that electronic structure of chemical elements (the **periodic table**) uses the fact that electrons can have 2 different spins. Hopefully you've heard about that in your lessons of chemistry.

Experiment on Quantum Eraser:

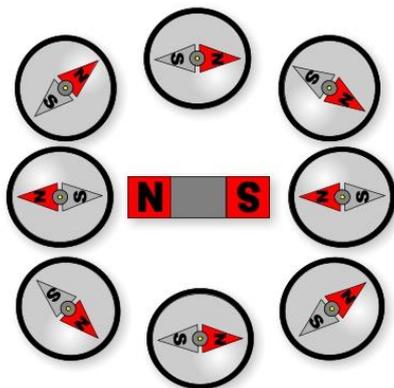
If you want to know more about Quantum eraser you may read more and make an interesting experiment proposed here:

<http://www.arturekert.org/miscellaneous/quantum-eraser.pdf>

4 Application: Magnetic Resonance Imaging (MRI)

As mentioned above, spins behave like small magnets. What do you see when you walk around with a compass?

.....
Explain why this happens:
.....



One of the properties of magnets is that they align with the magnetic field that surrounds them. This is why the needle of a compass always point to the north. It follows the direction of the magnetic field created by the earth.

Now, in the human body there is quite a lot of water. Each water molecule H_2O consists of 2 hydrogen atoms. Within each hydrogen H there is a proton with a net spin.



When undergoing an MRI scan (Magnetic Resonance Imaging) the patient is put into a cylindrical shaft. Within the casing of this cylinder there are strong magnets.

The spins in the body *tend to align with the magnetic field of the magnets in the machine*. **Contrary to the needle of a compass, the spins in the hydrogen atom can only point in 2 "directions": up or down.**

The spins *in their turn* also *create a magnetic field which can be measured by the machine!* Let us see how it works.

When the patient is inside the MRI scanner in fact *a changing magnetic field is added* which causes the spins in the body to *follow* the changing magnetic field. How many of the spins actually follow the changing magnetic field depends on the frequency at which it changes.

At **one specific frequency** the number of **changing spins** will be **maximal**. This phenomenon is called **nuclear magnetic resonance** (which is why such a scanner is also known as an NMR scanner).

You can try this out yourself with the following applet:

<http://phet.colorado.edu/en/simulation/mri>

The *change* in direction of the spins themselves can be measured because *the magnetic field created by them* also *changes* as they change direction. When the changing magnetic field of the machine *is turned off* again the *spins will go back* to being aligned with the

6 Concepts in Learning Station IX

Complete by adding the missing concepts

Classical concepts

Objects with spin behave like small magnets: particles with spin are deflected in a magnetic field. The amount by which they are deflected depends on

Quantum concepts

Spin is quantized: objects have and, in each dimension.

Each measurement of the spin in one particular direction acts like a "quantum eraser": the information gained in a previous measurement of spin in another direction is "erased".